

Chemistry Chapter 13 States Of Matter

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For this chapter, you will need to know about kinetic energy. Kinetic energy is what is created when an object moves. Kinetic theory also states that all matter consists of particles that are constantly in motion. Therefore, when in a closed container, all these moving particles that make up, well, everything will have to hit each other.

Chapter 13: States of Matter - Chemistry by Anna

Chemistry - Chp 13 - States of Matter 1. Chapter 13“States of Matter” 2. Section 13.1 The Nature of Gases OBJECTIVES: Describe the assumptions of the “kinetic theory” as it applies to gases. 3. Section 13.1 The Nature of Gases OBJECTIVES: Interpretgas pressure in terms of kinetic theory. 4.

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237 States of Matter Chapter 13 - States of Matter - 13.2 The Nature of Liquids - 13.2 Lesson Check - Page 430: 17 Answer liquid particles have attraction forces between them so they can't spread far away from each other so they have a definite volume.

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Chapter 12: states of matter Gases expand, diffuse, exert pressure, and can be compressed because they are in a low-density state consisting of tiny constantly-moving particles. Intermolecular forces, including dispersion forces, dipole-dipole forces, and hydrogen bonds, determine a substance's state at a given temperature.

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